

# Carbon Dioxide-Catalyzed Stereoselective Cyanation Reaction

Tamal Roy,<sup>†</sup> Myungjo J. Kim,<sup>‡,§,||</sup> Yang Yang,<sup>†,||</sup> Suyeon Kim,<sup>‡,§</sup> Gyumin Kang,<sup>‡,§</sup> Xinyi Ren,<sup>†</sup> Anders Kadziola,<sup>†</sup> Hee-Yoon Lee,<sup>\*,‡,§</sup> Mu-Hyun Baik,<sup>\*,‡,§</sup> and Ji-Woong Lee<sup>\*,†,§</sup>

<sup>†</sup>Department of Chemistry, University of Copenhagen, Universitetsparken 5, Copenhagen Ø 2100, Denmark

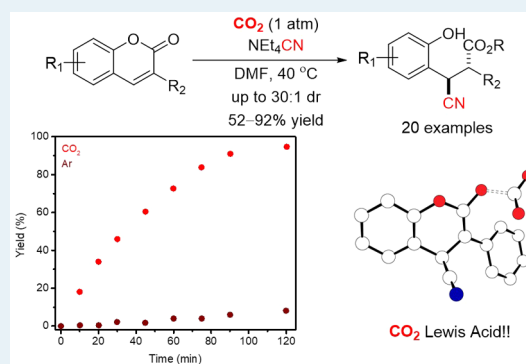
<sup>‡</sup>Department of Chemistry, Korea Advanced Institute of Science and Technology (KAIST), Daejeon 34141, Korea

<sup>§</sup>Center for Catalytic Hydrocarbon Functionalizations, Institute for Basic Science (IBS), Daejeon 34141, Korea

## Supporting Information

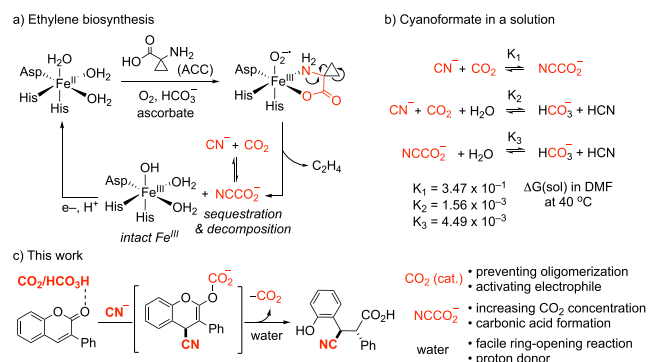
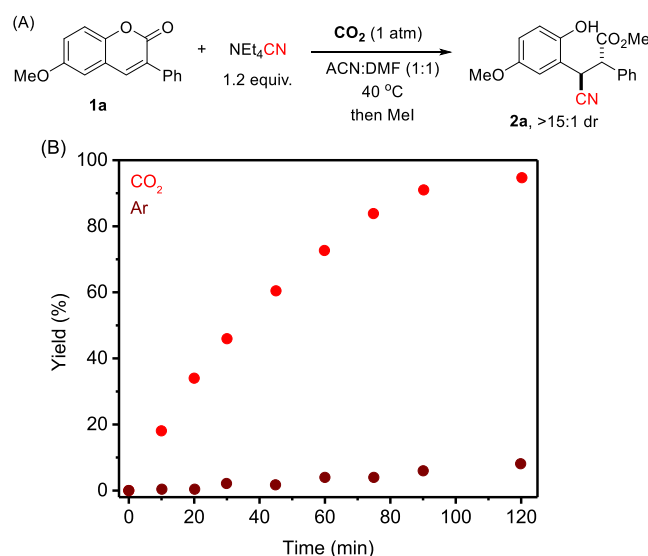
**ABSTRACT:** We report a Michael-type cyanation reaction of coumarins by using CO<sub>2</sub> as a catalyst. The delivery of the nucleophilic cyanide was realized by catalytic amounts of CO<sub>2</sub>, which forms cyanofornate and bicarbonate in the presence of water. Under ambient conditions, CO<sub>2</sub>-catalyzed reactions afforded high chemo- and diastereoselectivity of β-nitrile carbonyls, whereas only low reactivities were observed under argon or N<sub>2</sub>. Computational and experimental data suggest the catalytic role of CO<sub>2</sub>, which functions as a Lewis acid, and a protecting group to mask the reactivity of the product, suppressing byproducts and polymerization. The utility of this convenient method was demonstrated by preparing biologically relevant heterocyclic compounds with ease.

**KEYWORDS:** carbon dioxide, cyanation, cyanofornate, DFT calculation, coumarins



The cyanofornate ion (NC–CO<sub>2</sub><sup>−</sup>) is involved in the biological synthesis of ethylene at an Fe-containing enzyme (Figure 1a).<sup>1,2</sup> The formation of the cyanofornate ion from 1-aminocyclopropane-1-carboxylic acid is thought to protect the catalytically active Fe(III) center, which may be poisoned by cyanide ligation.<sup>3–5</sup> Although it is thermodynamically unstable, cyanofornate formed by combining cyanide and CO<sub>2</sub> may present a new opportunity in synthesis as a convenient source of highly nucleophilic cyanides avoiding the direct use of toxic hydrogen cyanide. Various surrogates of

**Scheme 1. (A) Hydrocyanation and Ring-Opening Reaction of Coumarin 1a under CO<sub>2</sub> and (B) a Comparison of Reactivities of 1a under CO<sub>2</sub> (Red) and Argon (Brown) as a Function of Time**



**Figure 1.** (a) Effects of cyanofornate in preserving the Fe(III) active center in the ethylene biosynthesis.<sup>5</sup> (b) CCSD(T)-calculated equilibria and potential advantage of CO<sub>2</sub> dissolution in cyanide and water-containing solutions. (c) Conjugate cyanation of coumarins.

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Table 1. Optimization of the Reaction Conditions<sup>a</sup>

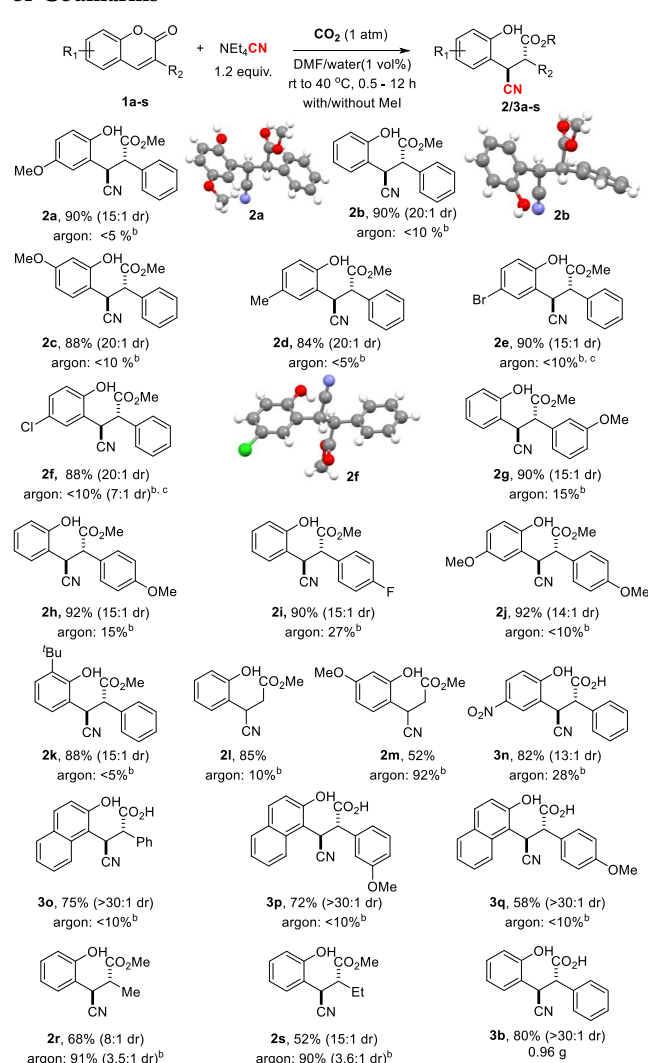
entry	deviation from standard reaction condition	yield (%) <sup>b</sup>
1	none	99 <sup>c</sup> (90) <sup>d</sup>
2	without additional water	90
3	argon or N <sub>2</sub> instead of CO <sub>2</sub>	<10
4	Presaturation with CO <sub>2</sub>	82
5	20–40 mol % of CO <sub>2</sub>	67–81%
6	MeCN as a solvent	99 <sup>e</sup>
7	at 25 °C	85
8	KCN instead of NEt <sub>4</sub> CN	n.d. <sup>f</sup>
9	3 equiv NaCN + NBu <sub>4</sub> Cl (10 mol %)	n.d.
10	3 equiv NaCN + 1.2 equiv NBu <sub>4</sub> Cl	72% <sup>g</sup>
11	Lewis, <sup>h</sup> Brønsted <sup>h</sup> acids instead of CO <sub>2</sub>	n.d.

<sup>a</sup>Standard reaction condition: substrate (**1b**, 0.2 mmol) and NEt<sub>4</sub>CN (0.24 mmol, 1.2 equiv) were dissolved in DMF/water (1 mL/0.01 mL) under CO<sub>2</sub> at 40 °C, and the yield was recorded after 1 h. MeI (0.6 mmol, 3 equiv) was added to isolate the methyl ester (**2b**) after workup. <sup>b</sup><sup>1</sup>H NMR yield using 1,3,5-trimethoxybenzene as an internal standard. <sup>c</sup>dr = 20:1. <sup>d</sup>Isolated yield (%). <sup>e</sup>dr = 6:1. <sup>f</sup>Not detected. <sup>g</sup>After 3 h. <sup>h</sup>Lewis acid and Brønsted acids (2 equiv): Ti(O<sup>i</sup>Pr)<sub>4</sub>, BF<sub>3</sub>·OEt<sub>2</sub>, Cu(OAc)<sub>2</sub>·H<sub>2</sub>O, FeCl<sub>3</sub>·6H<sub>2</sub>O, *m*-ClC<sub>6</sub>H<sub>4</sub>CO<sub>2</sub>H, NH<sub>4</sub>Cl, and *p*-nitrophenol were added instead of CO<sub>2</sub>. MeCN was used as a solvent.

HCN are known, but they are based on forming HCN *in situ*.<sup>6</sup> Recently, “shuttle” catalysis was introduced to deliver cyanide from less toxic alkyl nitriles.<sup>7–9</sup> Nonetheless, cyanides are indispensable as reagents for many reactions. Inspired by the ethylene biosynthesis, we envisaged that the cyanofornate may be useful in organic synthesis. Combining cyanide and CO<sub>2</sub> in solution affords cyanofornate,<sup>5</sup> as our experiments and calculations confirmed (Figure 1b), where the activities of the cyanofornate and bicarbonate remain uncovered.

To investigate this hypothesis (a synthetic application of cyanofornate in chemical reaction), we chose coumarin derivatives. There are currently no convenient procedures for accessing nitrile derivatives of coumarins, which represent densely functionalized synthons for organic synthesis. Owing to the divergent reactivity of alkyl nitriles to amines, carboxylic acids, aldehydes/ketones, and amides, cyanation is ubiquitous in natural products syntheses, pharmaceuticals, and polymers.<sup>10–12</sup> Hydrocyanation of unsaturated carbonyls is particularly interesting and results in β-cyano adducts, which can be further elaborated to γ-aminobutyric acids and 1,2-dicarboxylic acids.<sup>13</sup> However, 1,4-conjugate addition of hydrogen cyanide has been elusive due to competing 1,2-addition reactions.<sup>14,15</sup> Nagata and co-workers demonstrated that Lewis acids increased selectivity toward 1,4-addition, but an aluminum-based reagent was necessary starting from gaseous HCN.<sup>16,17</sup>

Herein, we report an operationally simple 1,4-cyanation reaction of activated substrates, namely coumarins, to afford β-cyano carboxylates using CO<sub>2</sub> and cyanide (Figure 1c). The presented reactions were highly chemo- and stereoselective even in the presence of water. A larger scale synthesis of α-aryl-β-cyano esters demonstrates the utility of the protocol for preparing heterocyclic compounds. In-depth computational analysis supports the proposed catalytic role of CO<sub>2</sub>,

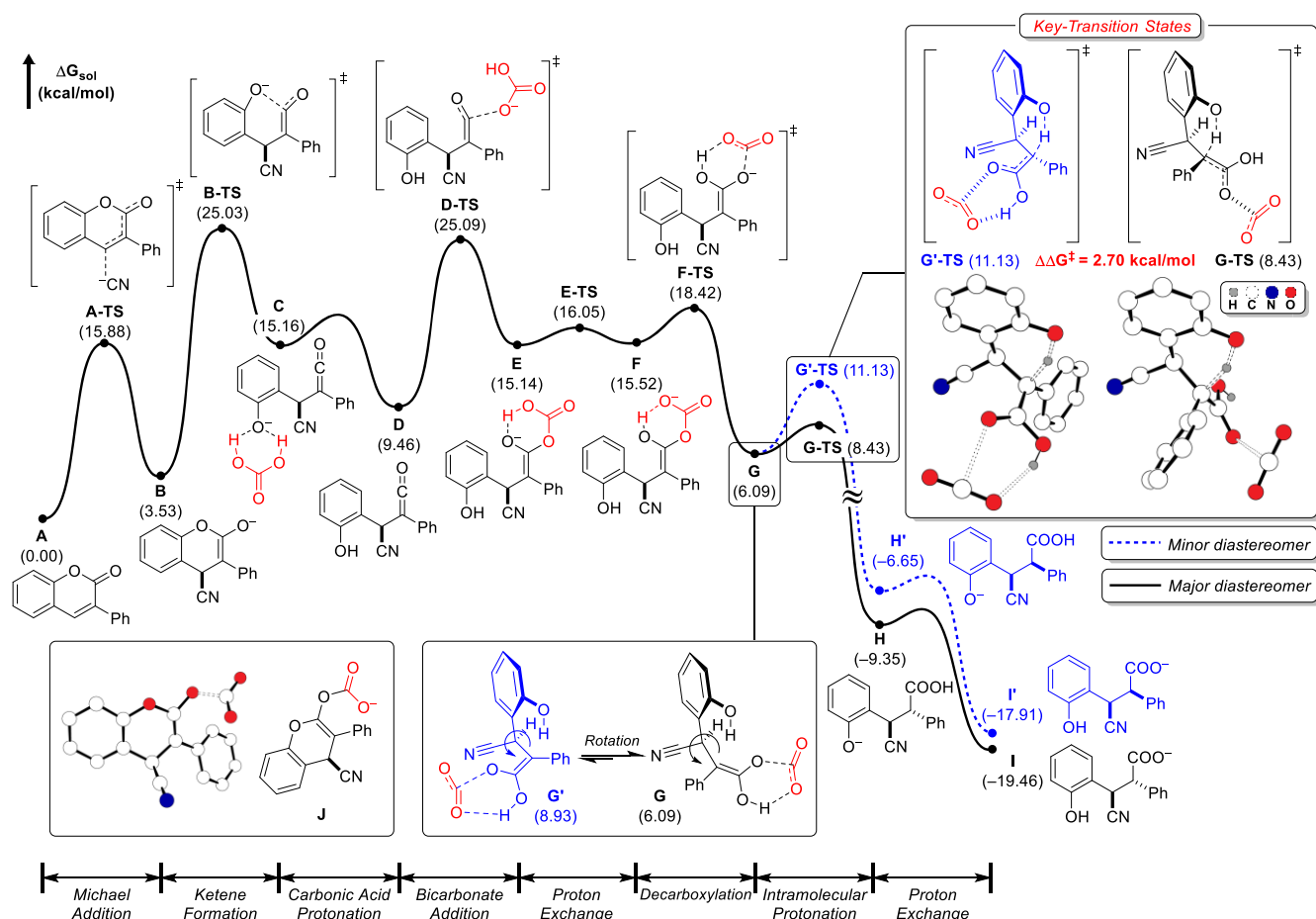
Scheme 2. Substrate Scope of the CO<sub>2</sub>-Catalyzed Cyanation of Coumarins<sup>a,b,c</sup>

<sup>a</sup>Reaction conditions: **1** (0.4 mmol), NEt<sub>4</sub>CN (0.48 mmol), DMF (2 or 4 mL), water (20 μL) and MeI (1.2 mmol). Products were isolated by crystallization or silica gel column chromatography. The control experiments under argon were conducted under identical conditions to the corresponding CO<sub>2</sub> experiments. <sup>b</sup>For control experiments under argon, conversion of starting materials was recorded by <sup>1</sup>H NMR using 1,3,5-trimethoxybenzene as an internal standard. <sup>c</sup>See Supporting Information section 11 for more details.

bicarbonate, and carbonic acid as Lewis and Brønsted acids to selectively activate electrophiles.

Recently, CO<sub>2</sub>-mediated organic reactions were reported<sup>18–20</sup> expanding its role beyond the well-known function as a cheap C1 source.<sup>21–28</sup> Fundamentally, adding CO<sub>2</sub> to cyanide and using it as a nucleophilic reagent is counterintuitive, as the nucleophilicity of cyanide is reduced upon forming cyanofornate.<sup>5</sup> Nevertheless, we were surprised to obtain high isolated yield (90%) of **2a** starting from coumarin **1a** with high diastereoselectivity (20:1 dr) under CO<sub>2</sub> atmosphere, whereas negligible reactivity was observed under N<sub>2</sub> or argon atmosphere (Scheme 1).

Table 1 summarizes the optimization of reaction conditions based on the methyl ester of the ring-opened product (**2b**) in the presence of an internal standard. The structure of **2b** was unambiguously confirmed by X-ray single crystal analysis



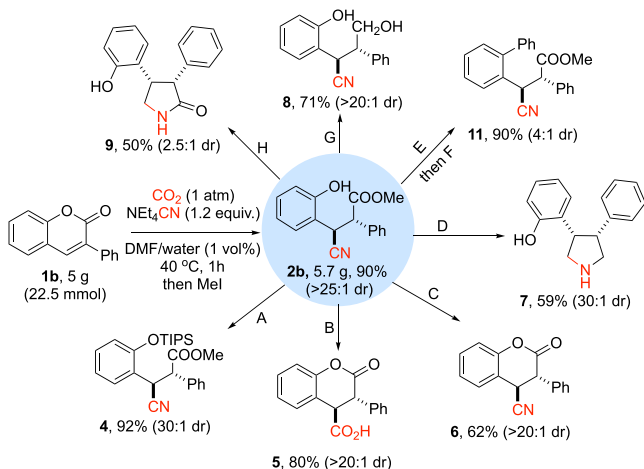
**Figure 2.** DFT calculated free energy profiles for the CO<sub>2</sub>-catalyzed cyanation of coumarin. The black trace represents the pathway to the major diastereomer. The blue trace represents the pathway to the minor diastereomer. All of the H atoms are removed for clarity at the ORTEP view.

(Scheme 2). The methylation of the  $\beta$ -cyano carboxylate intermediate was completely chemoselective without dimethylation.<sup>28</sup> A control experiment with <sup>13</sup>CO<sub>2</sub> strongly suggested no incorporation of atmospheric CO<sub>2</sub> in the final product (Figures S2 and S3). Reaction with D<sub>2</sub>O instead of H<sub>2</sub>O gave identical results (99% NMR yield, >15:1 dr) with >75% H–D exchange at both acidic protons on stereogenic carbons (Figure S6).

Interestingly, small amounts of water (~1 vol %) improved the reaction rate, yield, and selectivity toward the *syn*-product (entries 1 and 2), which may be due to water promoting the decomposition of cyanocarboxylate and faster release of cyanide.<sup>3</sup> When the same reaction was carried out under N<sub>2</sub> or argon atmosphere, only a trace amount (<10%) of **2b** was found with the majority of substrate remaining unreacted (entry 3). As a control experiment, a saturated solution was prepared by purging a solution of tetraethylammonium cyanide with CO<sub>2</sub>. The use of this solution afforded 82% yield of **2b** without additional gaseous CO<sub>2</sub>, confirming that the cyanation is feasible with dissolved CO<sub>2</sub> (entry 4). Moreover, we tested reactions with reduced CO<sub>2</sub> amounts (20–40 mol % with respect to coumarin **1b**) and found up to 81% conversion of **1b** to the desired product, suggesting a potential turnover of CO<sub>2</sub> (up to 3 TON, entry 5). Acetonitrile was a compatible solvent; however, the diastereoselectivity was reduced to 6:1, while lower conversion was observed at room temperature (entries 6 and 7). Various insoluble cyanide sources were inferior to soluble ammonium cyanide (entry 8 and Table S1).

Catalytic amounts of NBu<sub>4</sub>Cl (10 mol %) were employed to evaluate a potential phase-transfer catalysis, but excess stoichiometric amounts of NBu<sub>4</sub>Cl (1.2 equiv) were necessary to mediate the cyanation reaction (entries 9 and 10). Various Lewis and Brønsted acids, such as BF<sub>3</sub>·OEt<sub>2</sub>, Ti(O<sup>*i*</sup>Pr)<sub>4</sub>, Sc(OTf)<sub>3</sub>, Fe(III), Cu(II), NH<sub>4</sub>Cl, and HCl, with and without water showed negligible activity, implying a superior performance of CO<sub>2</sub> and water in promoting nucleophilic cyanide addition reactions. To the best of our knowledge, this new reaction protocol employing tetraethylammonium cyanide with CO<sub>2</sub> is the only method to selectively deliver cyanide for the conjugate addition reaction of coumarins.

With the optimized reaction conditions in our hands, we verified the generality of the cyanation reaction by varying the electronic properties of coumarins. The relative stereochemistry of products was unambiguously determined by X-ray single crystallography of products **2a**, **2b**, and **2f** (Scheme 2). High diastereoselectivity and isolated yield were obtained with corresponding  $\beta$ -cyano carboxylic esters/acids (**2a–2s**/**3n–q**) under CO<sub>2</sub> without sophisticated purification. The same reactions in the absence of CO<sub>2</sub> (under argon or N<sub>2</sub>), in general, afforded inferior conversion and diastereoselectivity of the cyanation products. Practical reaction conditions allowed us to smoothly convert coumarin **1b** to carboxylic acid product **3b** (0.96 g) after a single crystallization step. Various substituents (R<sub>2</sub>) on the 3-position were tolerated, including free carboxylic acid. For example, 3-carboxycoumarin, **1l**, was converted to the corresponding methyl ester (**2l**) after



**Figure 3.** Functional group transformations of  $\beta$ -cyano carboxylate derivatives **2b**. Reaction conditions: (A) **2b** (0.2 mmol), TIPSCl (1.5 equiv), imidazole (3 equiv), DMF (1 mL), rt, 3 h; (B) **2b** (0.2 mmol), 10 M HCl (2 mL), 70 °C, 16 h; (C) **2b** (0.2 mmol), Amberlyst-15 (50 mg), toluene (3 mL), 100 °C, 16 h; (D) **2b** (1 mmol),  $\text{BH}_3\cdot\text{Me}_2\text{S}$  (2 mmol), THF (10 mL), reflux, 10 h; (E) **2b** (1.78 mmol), pyridine (2 equiv, 3.56 mmol), triflic anhydride (1.2 equiv, 2.14 mmol), 0 °C to rt, 2 h, triflate protected phenol **10** was isolated in 93% yield (30:1 dr); (F) **10** (0.2 mmol)  $\text{PhB(OH)}_2$  (1.1 equiv, 0.22 mmol),  $\text{Pd(PPh}_3)_4$  (0.05 equiv, 0.01 mmol),  $\text{NEt}_3$  (2 equiv, 0.4 mmol), toluene/ethanol (3:1, 3 mL), microwave, 120 °C, 30 min; (G) **2b** (1 mmol),  $\text{LiBH}_4$  (2 M in THF, 2 equiv, 2 mmol), THF (10 mL), 50 °C, 2 h; (H) (i) **2b** (1 mmol), Pd/C (5 mol %), HCl (0.05 mL), MeOH (5 mL),  $\text{H}_2$  (1 atm), 48 h, (ii)  $\text{NEt}_3$  (1.2 equiv), toluene, reflux, 12 h.

decarboxylation. Moreover, we observed increased diastereoselectivity for products **2f**, **2r**, and **2s** under  $\text{CO}_2$  (>15:1 dr) compared to reactions under argon, where higher conversion was observed presumably due to stronger background reactions of alkyl substrates (see the [Supporting Information](#) for more details). It is worth noting here that cyanation reactions under argon generated unidentifiable byproducts including oligomers, which hampered the isolation of the product.

Although crystallization might improve the diastereoselectivity of the reaction, we ascribed the origin of high stereoselectivity (up to 30:1 dr) under the  $\text{CO}_2$  enriched solution to the presence of water and  $\text{CO}_2$ . As mentioned earlier, a  $\text{D}_2\text{O}$  experiment confirmed that two acidic protons are exchangeable, indicating potential racemization.<sup>29</sup> A set of control experiments with independently prepared cyanated product (**6**) subjected to the reaction conditions showed a facile elimination reaction to the starting material, confirming low barriers for the cyanation reaction, as confirmed by DFT calculations, summarized in [Figures 2](#) and [S7](#). In our DFT calculations, various reactivities of the enolate **B** have been considered, and we obtained intermediate **J** where  $\text{CO}_2$  is bound to the exocyclic oxygen atom of the enolate. This  $\text{CO}_2$  is significantly bent at an angle of 139° and the distance between its central carbon and the enolate oxygen was 1.631 Å (Mayer bond order = 0.58), implying that the  $\text{CO}_2$  acts as a Lewis acid. The resulting enolate **B** traverses the transition state **B-TS** at 25.0 kcal/mol. The ring opening reaction affords ketene intermediate **C**, which explains the observed reactivity with nucleophilic reagents, such as MeOH,  $i$ PrOH, and pyrrolidine, to give the corresponding esters and the amide ([Table S6](#)).

Several mechanistic scenarios involving the ketene were examined to shed light on the diastereoselective protonation, as detailed in the [Supporting Information](#). As shown in [Figure 2](#), the carbonate addition to the  $\text{C}=\text{O}$  bond is most plausible traversing **D-TS** at 25.1 kcal/mol to afford intermediate **E**, which quickly rearranges to form the proton exchange product **F**. On the basis of the experimental observations, the involvement of  $\text{CO}_2$  and related equilibria is plausible at the diastereoselective protonation step. With carbon dioxide not fully dissociated, the decarboxylation step is estimated to be downhill in energy by 9.4 kcal/mol. Notably, two major rotational isomers (**G** and **G'**) were identified to undergo facile intramolecular protonation with the barrier of only 2.3 kcal/mol, in the presence of weakly bound  $\text{CO}_2$ . The energy difference between the rotamers were maintained for the corresponding transition states leading to a differentiation for the *si*- and *re*-face protonation. The observed transition states **G-TS** and **G'-TS** showed a remarkable energy difference (2.7 kcal/mol), consistent with the experimental diastereoselectivity. In the absence of  $\text{CO}_2$  binding, as illustrated in [Supporting Information](#), the gap between two proton transfer barriers was reduced significantly, which results in the reduced diastereoselectivity, confirmed by experiments.

The utility of this highly stereoselective processes was demonstrated in a larger scale reaction (22.5 mmol) of **1b**. Owing to the high crystallinity of the product, methyl ester **2b** and carboxylic acid **3b** were precipitated in high yield and selectivity (80–90%, >25:1 dr, [Figure 3](#)). The employed ammonium cation can be recycled in high purity as tetraethylammonium iodide after methylation (see section 12, [Supporting Information](#)), avoiding HCN generation.<sup>30,31</sup>

Obtained product **2a** was subjected to functional group transformations to demonstrate the application potential of cyanated coumarins ([Figure 3](#)).<sup>32</sup> The phenolic  $-\text{OH}$  group on **2a** was protected to triisopropylsilyl ether **4** in high yield and diastereoselectivity. The hydrolysis of nitrile groups (**2b** and **3b**) under acidic conditions resulted in the same cyclic 1,4-dicarbonyl **5** with good yield and diastereoselectivity (80%, 20:1 dr). Acid-catalyzed reaction afforded 4-cyanodihydrocoumarin derivative **6** (62% 20:1 dr) starting from methyl ester **2b**. A borane reduction of methyl ester **2b** smoothly afforded pyrrolidine derivative **7** (59%, 30:1 dr) upon reduction of the nitrile group to a primary amine and *in situ* cyclization. A selective reduction of methyl ester **2b** over the nitrile group was performed with  $\text{LiBH}_4$  to furnish  $\gamma$ -cyano alcohol **8** (71% yield, >20:1 dr). Biologically important five-membered lactam derivative **9** was synthesized via the reduction of nitrile functional group with  $\text{H}_2$  over Pd/C, followed by cyclization in toluene under reflux (50% yield in two steps, 2.5:1 dr). Further structural diversification in the *ortho*-position of the aromatic substituent was achieved via Suzuki coupling of corresponding triflate derivative **10** to produce biaryl derivative **11** in high isolated yield (90%) with slightly diminished dr (4:1) due to the basic reagents in the reaction.

In conclusion, a catalytic application of  $\text{CO}_2$  and its equilibria with water and cyanide (to form cyanohydrate) were demonstrated by a 1,4-conjugate cyanation reaction of coumarins. This new and practical synthetic methodology enabled us to access various heterocycles of biological relevance. The role of  $\text{CO}_2$  was confirmed by mechanistic studies and computational analysis, pinpointing the critical involvement of  $\text{CO}_2$  in the product determining step and while affecting the reaction rate.<sup>33</sup> The proposed mechanism showed



a potential asymmetric catalysis where CO<sub>2</sub> is involved in a stereoselective step.<sup>29</sup> We are currently investigating various (chiral) ammonium and metal salts to expand the scope of CO<sub>2</sub>-catalyzed organic transformations.

## ■ ASSOCIATED CONTENT

### ■ Supporting Information

The Supporting Information is available free of charge on the ACS Publications website at DOI: 10.1021/acscatal.9b01087.

Methods, procedures, and characterization data; tables of screening data, tested Lewis and Brønsted acids, effects of CO<sub>2</sub> on cyanation, and crystal data; NMR spectra; free energy profiles; XYZ coordinates; ORTEP structures; HPLC chromatogram (PDF)

## ■ AUTHOR INFORMATION

### Corresponding Authors

\*Hee-Yoon Lee. E-mail: leehy@kaist.ac.kr.

\*Mu-Hyun Baik. E-mail: mbaik2805@kaist.ac.kr.

\*Ji-Woong Lee. E-mail: jiwoong.lee@chem.ku.dk.

### ORCID

Tamal Roy: 0000-0001-6813-0539

Myungjo J. Kim: 0000-0002-7031-7332

Yang Yang: 0000-0002-7588-1653

Suyeon Kim: 0000-0002-7032-1147

Hee-Yoon Lee: 0000-0002-3558-2894

Mu-Hyun Baik: 0000-0002-8832-8187

Ji-Woong Lee: 0000-0001-6177-4569

### Author Contributions

<sup>†</sup>These authors contributed equally.

### Notes

The authors declare no competing financial interest.

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